**Chemistry Chapter 5 Quiz #1**

Name \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Period \_\_\_\_ Date \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

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| **Questions 1-10: Name the following ionic compounds** | **Questions 11-20: Write the formula for the following ionic compounds** |
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| 1. CsF \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ | 1. aluminum sulfide \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ |
| 1. CaCO3 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ | 1. barium nitride \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ |
| 1. CuI2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ | 1. magnesium oxide \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ |
| 1. Mg(OH)2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ | 1. beryllium cyanide \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ |
| 1. Na3P \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ | 1. zinc (IV) nitrite\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ |
| 1. KC2H3O2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ | 1. gold (I) sulfate \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ |
| 1. Zr(Cr2O7)2\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ | 1. cobalt (III) carbonate \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ |
| 1. Fe2(SO3)3 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ | 1. Sodium permanganate \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ |
| 1. Ni3PO4 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ | 1. Zinc (II) chromate \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ |
| 1. AgNO3 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ | 1. Vanadium (IV) phosphate \_\_\_\_\_\_\_\_\_\_\_\_\_\_ |

**Questions 21-30: Match the terms on the left with the descriptions on the right.**

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| 1. \_\_\_\_ ion | 1. the combination of a sodium cation and a chlorine anion |
| 1. \_\_\_\_\_ anion | 1. the outermost electrons filling the s and p orbitals of an atom. |
| 1. \_\_\_\_ cation | 1. formed by an arrangement of repeating positively and negatively charged units within a salt |
| 1. \_\_\_\_\_ table salt | 1. elements loose or gain electrons to become like these stable elements |
| 1. \_\_\_\_\_ crystal lattice | 1. an element that has gained or lost an electron |
| 1. \_\_\_\_\_ octet rule | 1. an ion with a negative charge |
| 1. \_\_\_\_ ionic bond | 1. the tendency of an atom to gain or lose electrons, so that its outer shell has eight valence electrons |
| 1. \_\_\_\_\_ polyatomic ion | 1. an ion with a positive charge |
| 1. \_\_\_\_\_ noble gases | 1. formed by the electrostatic attraction of cation and anions |
| 1. \_\_\_\_\_ valence electrons | 1. an ion made up of two or more atoms |