**Chemistry Chapter 5 Quiz #1**

Name \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Period \_\_\_\_ Date \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

|  |  |
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| **Questions 1-10: Name the following ionic compounds**  | **Questions 11-20: Write the formula for the following ionic compounds** |
|  |  |
| 1. CsF \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
 | 1. aluminum sulfide \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
 |
| 1. CaCO3 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
 | 1. barium nitride \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
 |
| 1. CuI2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
 | 1. magnesium oxide \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
 |
| 1. Mg(OH)2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
 | 1. beryllium cyanide \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
 |
| 1. Na3P \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
 | 1. zinc (IV) nitrite\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
 |
| 1. KC2H3O2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
 | 1. gold (I) sulfate \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
 |
| 1. Zr(Cr2O7)2\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
 | 1. cobalt (III) carbonate \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
 |
| 1. Fe2(SO3)3 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
 | 1. Sodium permanganate \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
 |
| 1. Ni3PO4 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
 | 1. Zinc (II) chromate \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
 |
| 1. AgNO3 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_
 | 1. Vanadium (IV) phosphate \_\_\_\_\_\_\_\_\_\_\_\_\_\_
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**Questions 21-30: Match the terms on the left with the descriptions on the right.**

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| 1. \_\_\_\_ ion
 | 1. the combination of a sodium cation and a chlorine anion
 |
| 1. \_\_\_\_\_ anion
 | 1. the outermost electrons filling the s and p orbitals of an atom.
 |
| 1. \_\_\_\_ cation
 | 1. formed by an arrangement of repeating positively and negatively charged units within a salt
 |
| 1. \_\_\_\_\_ table salt
 | 1. elements loose or gain electrons to become like these stable elements
 |
| 1. \_\_\_\_\_ crystal lattice
 | 1. an element that has gained or lost an electron
 |
| 1. \_\_\_\_\_ octet rule
 | 1. an ion with a negative charge
 |
| 1. \_\_\_\_ ionic bond
 | 1. the tendency of an atom to gain or lose electrons, so that its outer shell has eight valence electrons
 |
| 1. \_\_\_\_\_ polyatomic ion
 | 1. an ion with a positive charge
 |
| 1. \_\_\_\_\_ noble gases
 | 1. formed by the electrostatic attraction of cation and anions
 |
| 1. \_\_\_\_\_ valence electrons
 | 1. an ion made up of two or more atoms
 |